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# Characterization of stockpiled phosphogypsum waste in Santos basin, Brazil

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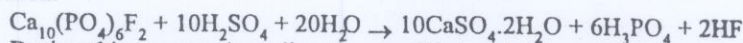
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**Abstract.** Phosphogypsum is a waste produced by the phosphate fertiliser industry. It is formed by precipitation during wet sulphuric acid processing of phosphate rocks. Although phosphogypsum is mainly calcium sulphate dihydrate, it contains elevated levels of impurities, which originate from the source phosphate rock used in the phosphoric acid production. Among these impurities, radionuclides from U-238 and Th-232 decay series, particularly Ra-226, Ra-228 and Pb-210, are of most concern due to their radiotoxicity. The phosphate fertiliser complex located in Santos Basin, Southwest Brazil, is responsible for the production of approximately 2000 tons of phosphogypsum per day. This phosphogypsum waste is stockpiled and presents a potential threat to the surrounding environment. This study aims to determine the activity concentration of Ra-226, Ra-228, Pb-210 and K-40 in the stockpiled phosphogypsum of the two main local producers. As a complementary study, trace and microelements (Ba, Co, Cr, Fe, Hf, Sb, Sc, Ta, Th, U, and rare earths Ce, Eu, La, Lu, Nd, Sm, Tb and Yb) were also determined by instrumental neutron activation analysis (INAA). Mean activity concentrations of  $401 \pm 108$  Bqkg<sup>-1</sup>,  $173 \pm 65$  Bqkg<sup>-1</sup> and  $389 \pm 106$  Bqkg<sup>-1</sup> were observed for Ra-226, Ra-228 and Pb-210 for one producer. The results obtained for the second producer were  $840 \pm 275$  Bqkg<sup>-1</sup>,  $225 \pm 32$  Bqkg<sup>-1</sup> and  $827 \pm 276$  Bqkg<sup>-1</sup> for Ra-226, Ra-228 and Pb-210, respectively. Phosphogypsum samples are enriched in rare-earth elements, specifically Ce, Eu, La, Nd, Sm, and Tb, and the elements Ba and Ta. The radiochemical and elemental characterisation of the phosphogypsum from industry A and C show that the stacks are quite homogeneous and mainly dependent upon the origin of the phosphatic rock used as raw material. This information is important and necessary for decision making on possible uses of this waste. The radionuclides Ra and Pb, which present high radiotoxicity, although present in the phosphogypsum, are not available for the surrounding environment.

## 1. INTRODUCTION

The presence of natural radionuclides in mineral ores and their redistribution in products and wastes has been well known. Among industrial wastes containing technologically enhanced naturally occurring radioactive materials (TENORM), excluding those generated by nuclear technology, of particular concern is phosphogypsum. This waste is produced world-wide in amounts of the order of 300 million tons per year, thus posing serious problems with its utilisation and safe disposal. In Brazil, several phosphate industries are responsible for the production of approximately 69 million tons of phosphogypsum waste. The two main industries located in Santos Basin, named "A" and "C", are responsible for the production of approximately 2000 tons of phosphogypsum per day [1]. No regulation in terms of radiation protection principles has been applied to these industries so far; as a consequence, members of the public may incur undue exposures.

The fundamental reaction in the phosphate fertiliser industry can be summarised by the following equation:



During this process the radioactive equilibrium between U, Th and their decay products is disrupted and the radionuclides migrate according to their solubility: uranium isotopes forms highly soluble compounds with the ion  $(\text{PO}_4)^{3-}$  while Ra and Th isotopes, Pb-210 and Po-210 concentrate into phosphogypsum [2].

While the ideal mole ratio between gypsum and phosphoric acid is 5:3 in the wet phosphoric acid process, the mass ratio is about 3:1; that means that the amount of gypsum produced correspond to three times the amount of acid. Phosphogypsum produced during the process is filtered off and pumped as slurry to nearby ponds, where it stays for a period sufficient to allow complete deposition. This waste is then moved to nearby storage areas, the so-called "gyp-stacks".

Although phosphogypsum is used in some countries for production of cement, construction materials and as a soil amendment in agricultural lands, its commercial use in Brazil is not well defined. This restriction is mainly due to the impurities contained in phosphogypsum, particularly Ra and Th isotopes.

PRODUÇÃO TÉCNICO CIENTÍFICA  
DO IPEN  
DEVOLVER NO BALCÃO DE  
EMPRÉSTIMO

Radiation - Colloques, v. 37, C1, p. 1307-1313, 2002.

ECORAD 2001

2001 TC - 8636 OK

2002 PB - 8637 OK



$^{210}\text{Pb}$  and  $^{210}\text{Po}$  derived from the U and Th series present in the phosphate rock. Besides the obvious waste of potentially valuable by-product, the main problem associated with phosphogypsum storage is the potential threat to the surrounding environment, specifically air and water resources in the vicinity of gypsum stacks. The most critical pathways between phosphogypsum and the public are, therefore, through inhalation of radon and through groundwater contamination.

This paper aims to determine the activity concentration of  $^{226}\text{Ra}$ ,  $^{228}\text{Ra}$ ,  $^{210}\text{Pb}$  and  $^{40}\text{K}$  in the stockpiled phosphogypsum of the two main local producers of phosphoric acid. Multielemental analysis of the phosphogypsum samples was also performed, by using instrumental neutron activation analysis (INAA) technique. Since one of the prime concerns about phosphogypsum storage is the potential for contamination of fresh water aquifers underlying the stacks, water from monitor wells located in the nearby was also collected for radionuclides determination.

## 2. MATERIAL AND METHODS

Phosphogypsum samples from stacks and water samples from monitor wells were collected in points depicted in Figure 1. In order to get information about radionuclides distribution throughout the deposit, cores (1 to 2.5 m in depth) were collected at different layers, from the surface to the bottom of the stacks. Stack from industry A was drilled in 19 different locations, whereas stacks from industry C were drilled in 10 different locations. The location of the sampling points were established by GPS (global positioning system) technique.

It is known that water exists in the crystalline lattice of both dihydrate and hemihydrate forms of phosphogypsum and there exist differences of up to 15.7% between the formula weights. Therefore, in order to get reliable results, the way the samples were dried and homogenised was extremely important, since all the radiochemical analysis performed are based upon gravimetric determinations. All the phosphogypsum samples studied were, therefore, prepared by drying at a temperature not greater than  $105^\circ\text{C}$  to constant mass, ground to a grain-size of less than  $250\ \mu\text{m}$  and finally stored in a desiccator at room temperature. Water samples were filtered and concentrated to a final volume of 100 ml.

Activity concentrations of  $^{228}\text{Ra}$ ,  $^{226}\text{Ra}$  and  $^{40}\text{K}$  were measured in samples of phosphogypsum and water by gamma spectrometry with a hyper-pure germanium detector, GEM-15200, from EG&G Ortec. The detector was calibrated using natural soil, rock and water spiked with radionuclides certified by Amersham. Samples were packed in  $100\ \text{cm}^3$  cans and sealed for about four weeks prior to the measurement in order to ensure that equilibrium has been reached between  $^{226}\text{Ra}$  and its decay products of short half-life. The  $^{226}\text{Ra}$  activities were determined by taking the mean activity of three separate photopeaks of its daughter nuclides:  $^{214}\text{Pb}$  at 295 keV and 352 keV, and  $^{214}\text{Bi}$  at 609 keV. The  $^{228}\text{Ra}$  content of the samples was determined by measuring the intensities of the 911 keV and 968 keV gamma-ray peaks from  $^{228}\text{Ac}$ . The potassium content was determined from the 1460 keV gamma-ray peak of  $^{40}\text{K}$ . Typical lower limits of detection for gamma spectrometry were 0.17 Bq/kg for  $^{226}\text{Ra}$ , 0.59 Bq/kg for  $^{228}\text{Ra}$  and 4.2 Bq/kg for  $^{40}\text{K}$ , for a counting time of 50,000 seconds.

The concentration of  $^{210}\text{Pb}$  was carried out by measuring the activity of its low energy peak (47 keV) in homogenised samples in a coaxial germanium detector, EGNC 15-190-R, from Eurisys. Samples were packed in  $100\ \text{cm}^3$  cans and counted for 50,000 seconds. Self-absorption correction was applied since the attenuation for low energy gamma rays is highly dependent upon sample composition. The approach used was modified from that suggested by Cutshall et al. [3]. According to them, the selfabsorption equation may be written as:

$$A/O = \frac{\ln\left(\frac{T}{I}\right)}{\left(\frac{T}{I} - 1\right)}$$

Where O is the weight in kg, I are the beam intensity (T) is defined as the source containing  $^{210}\text{Po}$  measurement with t represents how each activities are then ob as follows:

Where A is the weight in kg, fabs is seconds and  $\gamma$  is the counting time of 50,000 seconds. For the multielemental analysis (samples) and about 100 mg weighed and sealed reference material was used. R1m nuclear reactor second, after 15-20 minutes gamma spectrometry was performed with a resolution of 0.88 keV by Omnigam Spectrometer. Hf, Sb, Sc, Ta, Th, neutron activation analysis.

## 3. RESULTS AND DISCUSSION

The results obtained from the analysis of industry C and in Ta of the phosphogypsum C6, different groups with characteristics indicated in the samples are probably evaluated, disregarding the phosphogypsum. Measured activities observed for  $^{226}\text{Ra}$ ,  $^{228}\text{Ra}$ ,  $^{210}\text{Pb}$   $232 \pm 41\ \text{Bqkg}^{-1}$  and  $^{40}\text{K}$  previous determination of the radioactivity present in the ore rock, which is



Where O is the attenuated sample output (count rate), A is the actual sample photon emission rate, T and I are the beam intensities transmitted through the sample and standard, respectively. The attenuated beam intensity (T) is defined as the difference between the count rate of each sample with and without an external source containing  $^{210}\text{Pb}$  centred on top of the sample container. The factor I was determined by making a measurement with the same external  $^{210}\text{Pb}$  source placed on top of the efficiency standard. This ratio thus represents how each sample compares to the efficiency standard in terms of self-absorption. The final  $^{210}\text{Pb}$  activities are then obtained by substituting the relative absorption factor into the activity calculation equation, as follows:

$$A(\text{Bq/kg}) = \frac{(C - Co).f_{abs}}{Wt.\eta.T.\gamma}$$

Where A is the sample activity, C and Co are the sample and background counts, Wt is the sample weight in kg,  $f_{abs}$  is the relative absorption factor,  $\eta$  is the efficiency in cps/dps, T is the counting time in seconds and  $\gamma$  is the peak intensity. The lower limit of detection obtained for  $^{210}\text{Pb}$  was 20.6 Bq/kg, for a counting time of 50,000 seconds.

For the multielemental analysis of samples, approximately 200 mg of phosphogypsum (duplicate samples) and about 150 mg of reference material (standard MR IWG-GIT Granite AC-E) were accurately weighed and sealed in pre-cleaned double polyethylene bags, for irradiation. Phosphogypsum samples and reference material were irradiated for 16 hours, under a thermal neutron flux of  $10^{12} \text{ n cm}^{-2} \text{ s}^{-1}$  in the IEA-R1m nuclear reactor at IPEN. Two series of counting were made: the first, after one week decay and the second, after 15-20 days. The counting time was 1.5 hours for each sample and reference material. Gamma spectrometry was performed using an Intertechnique hyperpure Ge detector and associated electronics, with a resolution of 0.88 keV and 1.90 keV for  $^{57}\text{Co}$  and  $^{60}\text{Co}$ , respectively. The analysis of the data was made by Omnigam Spectrum Analysis to identify the gamma-ray peaks. Trace and micro elements (Ba, Co, Cr, Fe, Hf, Sb, Sc, Ta, Th, U and rare earths Ce, Eu, La, Lu, Nd, Sm, Tb and Yb) were determined by instrumental neutron activation analysis.

### 3. RESULTS AND DISCUSSION

The results obtained for the activity concentration of  $^{226}\text{Ra}$ ,  $^{228}\text{Ra}$ ,  $^{210}\text{Pb}$  and  $^{40}\text{K}$  are presented in Table 1 for industry C and in Table 2 for industry A. Cluster analysis was performed in order to verify the homogeneity of the phosphogypsum samples, from stacks A and B (Figure 2 and 3, respectively). Samples of phosphogypsum C6, C8 and C10 from industry C and samples A5, A9 and A10 from industry A represents different groups with no similarities with the other analysed samples. Their composition and physical characteristics indicate that: a) these samples are a mixture of phosphogypsum, soil and clay or b) these samples are probably originated from phosphate rocks from other countries. Average concentrations were evaluated, disregarding these samples and considering only samples representative of Brazilian phosphogypsum. Mean activity concentrations of  $401 \pm 108 \text{ Bqkg}^{-1}$ ,  $173 \pm 65 \text{ Bqkg}^{-1}$  and  $389 \pm 106 \text{ Bqkg}^{-1}$  were observed for  $^{226}\text{Ra}$ ,  $^{228}\text{Ra}$  and  $^{210}\text{Pb}$  in industry C. The results obtained for producer A were  $845 \pm 266 \text{ Bqkg}^{-1}$ ,  $232 \pm 41 \text{ Bqkg}^{-1}$  and  $834 \pm 267 \text{ Bqkg}^{-1}$  for  $^{226}\text{Ra}$ ,  $^{228}\text{Ra}$  and  $^{210}\text{Pb}$ , respectively. These results agree well with previous determination of the same radionuclides in fresh phosphogypsum by Mazzilli et al [2], showing that the radioactivity present in the stockpiled phosphogypsum is mainly dependent on the amount supplied by the ore rock, which is different for the two industries.



Table 1: Activity concentrations of  $^{228}\text{Ra}$ ,  $^{226}\text{Ra}$ ,  $^{210}\text{Pb}$  and  $^{40}\text{K}$  in phosphate rock (PR) and phosphogypsum (PG) from producer C and soil ( $\text{Bq kg}^{-1}$ ).

Samples	Collecting date	$^{228}\text{Ra}$	$^{226}\text{Ra}$	$^{210}\text{Pb}$	$^{40}\text{K}$
PR C	06/19/2000	221 $\pm$ 28	427 $\pm$ 47	422 $\pm$ 29	<41
PG C1		130 $\pm$ 16	350 $\pm$ 42	353 $\pm$ 31	<52
C2		90 $\pm$ 12	594 $\pm$ 65	581 $\pm$ 35	<43
C3		229 $\pm$ 19	436 $\pm$ 62	422 $\pm$ 28	<61
C4		238 $\pm$ 21	307 $\pm$ 48	303 $\pm$ 15	<55
C5		214 $\pm$ 19	337 $\pm$ 42	318 $\pm$ 19	15 $\pm$ 1
C6*	09/27/2000	87 $\pm$ 7	56 $\pm$ 4	49 $\pm$ 3	<92
C7		96 $\pm$ 6	480 $\pm$ 36	464 $\pm$ 33	<95
C8*		<12	177 $\pm$ 14	159 $\pm$ 13	<88
C9		217 $\pm$ 19	303 $\pm$ 22	287 $\pm$ 17	<91
C10*		98 $\pm$ 7	59 $\pm$ 4	43 $\pm$ 4	<78
Mean $\pm$ standard deviation		173 $\pm$ 65	401 $\pm$ 108	389 $\pm$ 106	-
Soil**		62 $\pm$ 24	88 $\pm$ 31	-	454 $\pm$ 173

\* Results not considered in the evaluation of the arithmetic mean

\*\* Results for soil represent the arithmetic mean of 8 soil samples collected in the surrounding environment of industry A and C.

Table 2 Activity concentrations of  $^{228}\text{Ra}$ ,  $^{226}\text{Ra}$ ,  $^{210}\text{Pb}$  and  $^{40}\text{K}$  in water samples from the monitor wells ( $\text{Bq L}^{-1}$ ).

Sampling location	Collecting date	$^{228}\text{Ra}$	$^{226}\text{Ra}$	$^{210}\text{Pb}$	$^{40}\text{K}$
MW 02	06/20/2000	8.0 $\pm$ 0.4	1.7 $\pm$ 0.1	<1.6	8.1 $\pm$ 0.4
MW 03		4.5 $\pm$ 0.2	1.2 $\pm$ 0.1	2.0 $\pm$ 0.1	7.9 $\pm$ 0.5
MW 04		4.8 $\pm$ 0.2	1.2 $\pm$ 0.1	3.1 $\pm$ 0.2	<4.1
MW 05		6.7 $\pm$ 0.3	2.9 $\pm$ 0.2	<1.6	11 $\pm$ 1
MW02	09/28/2000	<0.9	1.1 $\pm$ 0.1	1.1 $\pm$ 0.1	<4.1
MW03		3.1 $\pm$ 0.1	1.2 $\pm$ 0.1	<1.6	<4.1
MW 04		1.4 $\pm$ 0.1	1.5 $\pm$ 0.1	<1.6	<4.1
Mean $\pm$ standard deviation		4.1 $\pm$ 2.6	1.5 $\pm$ 0.6	1.8 $\pm$ 1.4	-

In Table 3 are presented samples collected in the exceed 1.8  $\text{Bq L}^{-1}$ , where mean activity concentration the recommended radio mean activity concentration for gross-alpha activity for drinking water supplied four locations presenting potential release of radionuclides found concentrations in fluids in Florida.

Table 3 Activity concentrations

Samples
PR A
PG A1
A2
A3
A4
A5*
A6
A7
A8
A9*
A10*
A11
A12
A13
A14
A15
A16
A17
A18
A19
Mean $\pm$ s

\* Results not considered in

3) from producer C

<sup>40</sup> K
41
52
43
61
55
±1
92
95
88
91
<78
-
±173

of industry A and C

r wells

<sup>40</sup> K
8.1±0.4
7.9±0.5
<4.1
11±1
<4.1
<4.1
<4.1
-

In Table 3 are presented results obtained for the activity concentration of <sup>228</sup>Ra, <sup>226</sup>Ra and <sup>210</sup>Pb in water samples collected in the monitor wells depicted in Figure 1. <sup>226</sup>Ra and <sup>210</sup>Pb mean concentrations did not exceed 1.8 Bq L<sup>-1</sup>, whereas mean activity of 4.1 Bq L<sup>-1</sup> was observed for <sup>228</sup>Ra. The results obtained for the mean activity concentrations of <sup>228</sup>Ra and <sup>210</sup>Pb in the monitor wells are of the same order of magnitude than the recommended radioactivity drinking water standards for gross-beta activity (1 Bq L<sup>-1</sup>) [4]. For <sup>226</sup>Ra the mean activity concentration is 20 times higher than the recommended radioactivity drinking water standards for gross-alpha activity (0.1 Bq L<sup>-1</sup>) [4]. Oliveira et al [5] measuring the activity concentration of <sup>226</sup>Ra in drinking water supplies of São Paulo State, found concentrations ranging from <2.2 to 235 mBq L<sup>-1</sup>; with four locations presenting <sup>226</sup>Ra activity higher than 0.1 Bq L<sup>-1</sup>. Recent papers [6,7,8] are concerned with the potential release of radionuclides from phosphogypsum to the aquatic environment. Burnett and Elzerman [6] found concentrations ranging from 0.1 to 0.2 Bq L<sup>-1</sup> for <sup>226</sup>Ra and from 7 to 70 Bq L<sup>-1</sup> for <sup>210</sup>Pb, in stack fluids in Florida.

Table 3 Activity concentrations of <sup>228</sup>Ra, <sup>226</sup>Ra, <sup>210</sup>Pb and <sup>40</sup>K in phosphate rock (PR) and phosphogypsum (PG) from producer A (Bq kg<sup>-1</sup>).

Samples	Collecting date	<sup>228</sup> Ra	<sup>226</sup> Ra	<sup>210</sup> Pb	<sup>40</sup> K
PR A	06/20/2000	260±27	638±61	645±39	19±1
PG A1		280±27	773±72	761±45	<46
A2		238±26	762±61	748±53	<57
A3		215±19	849±78	864±69	22±2
A4		269±32	603±27	592±41	<53
A5*		139±12	282±24	267±16	288±15
A6		255±28	1013±84	1005±81	153±11
A7		219±23	935±87	917±55	28±2
A8		198±17	1251±132	1234±63	<57
A9*		96±11	82±7	77±4	460±32
A10*		87±10	59±8	68±4	285±16
A11		263±22	558±62	539±43	<53
A12	09/28/2000	202±16	1117±85	1105±88	<106
A13		246±27	538±39	552±47	<78
A14		218±19	871±62	853±43	<81
A15		214±26	473±29	438±28	370±26
A16		211±24	472±33	451±27	<75
A17		334±32	928±71	951±76	<80
A18		191±16	1101±75	1089±65	<70
A19		163±15	1282±115	1257±75	<104
Mean ± standard deviation		232±41	845±266	834±267	-

\* Results not considered in the evaluation of the arithmetic mean



Tables 4 and 5 show the results obtained for multielemental analysis of phosphogypsum and phosphate rock samples by INAA. The precision based in duplicate samples analysis was better than 10%. Here too, the results obtained for elemental concentration in samples A5, A9 and A10 (Table 5) and C6, C8 and C10 (Table 4) presented a different pattern, confirming that the raw material used in the manufacturing of this phosphoric acid came from other locations or that the phosphogypsum samples were mixed with soil and clay in the stacks. Phosphogypsum samples are enriched in rare-earth elements, specifically Ce, Eu, La, Nd, Sm, and Tb, and the elements Ba and Ta.

Table 4 Concentrations of elements determined by INAA in phosphate rock (PR) and phosphogypsum (PG) from producer C ( $\mu\text{g g}^{-1}$ ).

Sample	La	Sm	Nd	Eu	Tb	Ce	Lu	Yb	U
PR C	2284 ± 120	264 ± 9	2026 ± 203	67 ± 4	12 ± 1	4342 ± 141	0.37 ± 0.02	10 ± 1	41 ± 4
PG C1	155 ± 9	23 ± 1	147 ± 10	5.2 ± 0.3	2.0 ± 0.1	307 ± 13	0.48 ± 0.02	2.6 ± 0.1	4.9 ± 0.6
C2	349 ± 19	49 ± 2	358 ± 25	13 ± 1	3.9 ± 0.2	639 ± 25	0.41 ± 0.02	3.3 ± 0.1	3.5 ± 0.1
C3	1300 ± 70	171 ± 6	1285 ± 88	36 ± 2	7.3 ± 0.3	2574 ± 99	0.11 ± 0.01	4.4 ± 0.2	2.9 ± 0.3
C4	1632 ± 88	149 ± 5	1722 ± 174	47 ± 3	10.1 ± 0.5	3387 ± 129	0.28 ± 0.01	4.3 ± 0.2	3.4 ± 0.3
C5	2373 ± 125	274 ± 9	1376 ± 139	46 ± 3	8.2 ± 1.1	3349 ± 109	0.36 ± 0.06	9.8 ± 0.6	3.0 ± 0.1

Sample	Ba	Co	Cr	Sb	Fe	HI	Sc	Ta	Th
PR C	15528 ± 5287	5.1 ± 0.9	34 ± 14	--	11892 ± 356	5.9 ± 0.6	29 ± 1	21 ± 2	100 ± 6
PG C1	964 ± 94	1.2 ± 0.6	23 ± 8	0.7 ± 0.1	21387 ± 707	3.3 ± 0.2	8 ± 4	1.2 ± 0.1	17 ± 1
C2	1525 ± 149	0.14 ± 0.07	9 ± 3	0.5 ± 0.1	565 ± 19	3.4 ± 0.3	0.5 ± 0.2	1.1 ± 0.1	15 ± 1
C3	2915 ± 283	0.10 ± 0.05	29 ± 10	1.1 ± 0.2	3996 ± 149	4.1 ± 0.2	4 ± 2	11 ± 1	43 ± 2
C4	2839 ± 283	0.4 ± 0.2	20 ± 7	1.1 ± 0.2	12447 ± 508	2.3 ± 0.1	3 ± 2	6 ± 1	66 ± 3
C5	5010 ± 1706	0.9 ± 0.6	29 ± 12	1.2 ± 0.2	2004 ± 336	1.2 ± 0.4	5.3 ± 2.5	6.9 ± 1.4	70 ± 3

Table 5 Concentrations of elements determined by INAA in phosphate rock (PR) and phosphogypsum (PG) from producer A ( $\mu\text{g g}^{-1}$ ).

Sample	La	Sm	Nd	Eu	Tb	Ce	Lu	Yb	U
PRA	2002 ± 106	265 ± 9	1920 ± 193	65 ± 4	10.8 ± 0.9	4277 ± 140	0.59 ± 0.04	7.9 ± 0.4	53 ± 1
PG A1	1602 ± 85	208 ± 7	1093 ± 105	37 ± 2	7.0 ± 0.6	2333 ± 76	0.29 ± 0.02	6.5 ± 0.3	3.2 ± 0.2
A2	1221 ± 65	172 ± 6	1181 ± 119	41 ± 2	9.0 ± 0.8	2628 ± 86	0.20 ± 0.01	4.8 ± 0.3	3.4 ± 0.3
A3	807 ± 41	113 ± 4	1399 ± 98	43 ± 3	5.9 ± 0.3	2456 ± 80	0.17 ± 0.01	2.5 ± 0.1	2.9 ± 0.3
A4	974 ± 36	154 ± 4	1203 ± 82	41 ± 5	7.1 ± 0.2	2357 ± 54	0.22 ± 0.01	3.4 ± 0.2	5.6 ± 0.4
A5	59 ± 3	8.9 ± 0.3	75 ± 9	2.0 ± 0.1	0.7 ± 0.1	1236 ± 41	0.28 ± 0.02	2.1 ± 0.1	4.2 ± 0.4
A6	1625 ± 83	158 ± 6	1409 ± 98	39 ± 2	7.5 ± 0.3	2877 ± 94	0.25 ± 0.01	5.0 ± 0.2	3.3 ± 0.2
A7	1256 ± 47	128 ± 3	1087 ± 54	41 ± 5	5.4 ± 0.2	2523 ± 58	0.15 ± 0.01	2.5 ± 0.1	4.3 ± 0.1
A8	939 ± 47	129 ± 5	1511 ± 159	34 ± 2	7.1 ± 0.4	1998 ± 63	0.18 ± 0.01	3.7 ± 0.2	2.6 ± 0.1
A9	38 ± 2	6.7 ± 0.2	35 ± 4	0.94 ± 0.05	1.2 ± 0.1	83 ± 3	0.55 ± 0.03	3.6 ± 0.3	3.2 ± 0.4
A10	36 ± 2	6.4 ± 0.2	30 ± 2	0.5 ± 0.8	0.7 ± 0.1	76 ± 3	0.34 ± 0.02	2.1 ± 0.2	1.6 ± 0.2
A11	1311 ± 66	170 ± 6	1378 ± 104	16 ± 1	9.2 ± 0.5	2588 ± 82	0.29 ± 0.01	5.4 ± 0.2	2.8 ± 0.3

Sample	Ba	Co	Cr	Sb	Fe	HI	Sc	Ta	Th
PRA	1207 ± 128	4.7 ± 2.9	30 ± 12	--	2156 ± 44	2.3 ± 0.2	21 ± 10	24 ± 3	77 ± 4
PG A1	1810 ± 618	0.8 ± 0.5	24 ± 10	--	2372 ± 78	2.3 ± 0.2	4 ± 2	8.9 ± 0.8	70 ± 3
A2	10433 ± 3552	1.1 ± 0.7	22 ± 9	0.3 ± 0.5	788 ± 29	1.6 ± 0.1	3 ± 1	7.2 ± 0.8	72 ± 3
A3	7046 ± 654	5.2 ± 3.2	22 ± 8	0.6 ± 0.1	2156 ± 44	2.2 ± 0.1	3 ± 1	9.1 ± 0.6	62 ± 3
A4	8192 ± 761	3.4 ± 1.7	26 ± 7	0.8 ± 0.1	1347 ± 30	2.3 ± 0.1	3 ± 1	7.2 ± 0.3	64 ± 2
A5	683 ± 73	23 ± 11	39 ± 14	0.7 ± 0.1	37419 ± 724	4.5 ± 0.4	17 ± 8	1.4 ± 0.1	12 ± 1
A6	4992 ± 465	6.1 ± 4.3	28 ± 10	0.8 ± 0.2	2285 ± 37	1.3 ± 0.1	4 ± 2	8.8 ± 0.6	69 ± 3
A7	3744 ± 248	1.3 ± 0.6	19 ± 7	0.8 ± 0.1	128 ± 3	1.6 ± 0.1	1.2 ± 0.5	2.3 ± 0.1	85 ± 2
A8	7622 ± 743	7.3 ± 3.6	175 ± 62	0.4 ± 0.1	1368 ± 64	2.3 ± 0.1	4 ± 2	--	51 ± 2
A9	430 ± 44	17 ± 12	42 ± 15	0.6 ± 0.5	43071 ± 555	8.9 ± 0.7	26 ± 12	1.4 ± 0.1	13 ± 1
A10	299 ± 31	5.9 ± 2.1	27 ± 10	0.4 ± 0.1	34264 ± 445	7.1 ± 0.6	15 ± 7	1.1 ± 0.1	17 ± 1
A11	1898 ± 188	2.1 ± 0.4	19 ± 7	1.0 ± 0.2	563 ± 12	8.3 ± 0.4	4 ± 2	6.5 ± 0.5	65 ± 3

#### 4. CONCLUSIONS

Phosphogypsum sample specifically Ce, Eu, La. characterization of the p and mainly dependent t important and necessary which present high ra surrounding environmer not relevant.

#### Acknowledgments

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4. CONCLUSIONS

Phosphogypsum samples from stacks are enriched in <sup>226</sup>Ra, <sup>228</sup>Ra and <sup>210</sup>Pb, in rare-earth elements, specifically Ce, Eu, La, Nd, Sm, and Tb, and in elements Ba and Ta. The radiochemical and elemental characterization of the phosphogypsum from industry A and C show that the stacks are quite homogeneous and mainly dependent upon the origin of the phosphatic rock used as raw material. This information is important and necessary for decision making on possible uses of this waste. The radionuclides Ra and Pb, which present high radiotoxicity, although present in the phosphogypsum, are not available for the surrounding environment. Indeed, the analysis of the water from the monitor wells gave results considered not relevant.

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and phosphate  
%. Here too, the  
C6, C8 and C10  
facturing of this  
d with soil and  
Ce, Eu, La, Nd,

ducer C (µg g<sup>-1</sup>).

Yb	U
10 ± 1	41 ± 4
2.6 ± 0.1	4.9 ± 0.6
3.3 ± 0.1	3.5 ± 0.1
4.4 ± 0.2	2.9 ± 0.3
4.3 ± 0.2	3.4 ± 0.3
9.8 ± 0.6	3.0 ± 0.1

Ta	Tb
21 ± 2	100 ± 6
1.2 ± 0.1	17 ± 1
1.1 ± 0.1	15 ± 1
11 ± 1	43 ± 2
6 ± 1	66 ± 3
6.9 ± 1.4	70 ± 3

ducer A (µg g<sup>-1</sup>).

Yb	U
7.9 ± 0.4	53 ± 1
6.5 ± 0.3	3.2 ± 0.2
4.8 ± 0.3	3.4 ± 0.3
2.5 ± 0.1	2.9 ± 0.3
3.4 ± 0.2	5.6 ± 0.4
2.1 ± 0.1	4.2 ± 0.4
5.0 ± 0.2	3.3 ± 0.2
2.5 ± 0.1	4.3 ± 0.1
3.7 ± 0.2	2.6 ± 0.1
3.6 ± 0.3	3.2 ± 0.4
2.1 ± 0.2	1.6 ± 0.2
5.4 ± 0.2	2.8 ± 0.3
Ta	Tb
24 ± 3	77 ± 4
8.9 ± 0.8	70 ± 3
7.2 ± 0.8	72 ± 3
9.1 ± 0.6	62 ± 3
7.2 ± 0.3	64 ± 2
1.4 ± 0.1	12 ± 1
8.8 ± 0.6	69 ± 3
2.3 ± 0.1	85 ± 2
-	51 ± 2
1.4 ± 0.1	13 ± 1
1.1 ± 0.1	17 ± 1
6.5 ± 0.5	65 ± 3